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Two equilibria must be considered: $\text{CuO (s)} + \text{H}_2 \text{ (g)} \rightleftharpoons \text{H}_2\text{O (l)} + \text{Cu}$ at equilibrium -
 $n_{\text{O}_2} = 2 - \text{Cu (s)} + \frac{1}{2} \text{O}_2 \text{ (g)} \rightleftharpoons \text{CuO}$ 1 at equilibrium - $n_{\text{O}_2} = 2 - 2$ In the final state,

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only O₂ is present. Therefore $n_{\text{O}_2} = n_{\text{O}}/2$. In an excess of O₂, all the 1 copper is oxidized. Therefore $n_{\text{Cu}} = 0$ or $n_{\text{Cu}} = n_{\text{O}}/2$.

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Thomas Engel has taught chemistry for more than 20 years at the University of Washington, where he is currently Professor of Chemistry and Associate Chair for the Undergraduate Program. Professor Engel received his bachelor's and master's degrees in chemistry from the Johns Hopkins University, and his Ph.D. in chemistry from the University of Chicago.

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