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CHAPTER 9 REVIEW Stoichiometry SECTION 3
PROBLEMS Write the answer on the line to the left.

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Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N₂ are mixed with 12.0 mol of H₂ according to the following equation: $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$

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Describe the relationship between Avogadro's number and a mole of a substance. A mole is defined as the amount of a substance in 6.022×10^{23} (Avogadro's number) particles.

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Review Problems □ Do section review problems #1
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chemical reactions. $2 \text{ HgO} \rightarrow 2 \text{ Hg} + \text{O}_2$ 2- a) 2 mol

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SHORT ANSWER Answer the following questions in the space provided. 1. a.

6 Chemical Bonding

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