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Read the
periodic table
from top left to
bottom right.

The elements are
ordered by their

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atomic numbers, which increase as you move across and down the periodic table. The atomic number is how many protons the element's atom possesses. You'll also notice that each element's atomic mass increases

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Chapter 5 Nonmetal V
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Hall Inc. Design
of the Periodic
Table Key

Concept:

Important
information
about an element
is given in each
square of the
periodic table:
Its atomic
number, chemical

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symbol, name, and atomic mass. Decide if each of the

underlined terms is used correctly. If it is, write correct after the sentence.

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Periodic Table

The Periodic

Table can

provide a wealth

of information

if it is

interpreted

properly. Click

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*The Arrangement
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Q5 Based on a
typical modern
version of the

Periodic Table

(a) In order of
atomic (proton)
number. (b) Due
to isotopic
masses, the
relative atomic
mass does go
'down'

occasionally

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(there is no obvious 'nuclear' rule that applies). BUT chemically Te is S and Se etc. and I is like Cl and Br etc.

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The periodic table is useful for understanding atomic properties that show periodic trends. One such property is the atomic radius (Figure 2.7. 5). The atomic radius is defined as one-

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half the distance between the nuclei of identical atoms that are bonded together. The units for atomic radii are picometers, equal to 10^{-12} meters.

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