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Electroplating - Wikipedia

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Harb = 10 - 14 Chapter 2 Problem 2.19 1/1 Determine the solubility product K_{sp} for $PbSO_4$. The desired equilibrium is $PbSO_4 \rightleftharpoons Pb^{2+} + SO_4^{2-}$. We can write this as the sum of two electrochemical equations $PbSO_4 + 2e^- \rightleftharpoons Pb + SO_4^{2-}$ and $Pb \rightleftharpoons Pb^{2+} + 2e^-$.

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