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Unit 1 chapter 12 stoichiometry

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Mole Ratio Practice Problems

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### Chapter 12 . Stoichiometry Notes

Packet Big Picture Ideas: The identity of the reactants helps scientists to predict the products in a chemical reaction. Quantitative relationships exist with all chemical reactions that allow scientists to predict amounts of products formed, reactants consumed, and percent yield based on theoretical maximum.

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## ~~CHAPTER 11: STOICHIOMETRY~~

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In Example 12.2.1 and Example 12.2.2, the identity of the limiting reactant has been apparent:  $[\text{Au}(\text{CN})_2]^-$ ,  $\text{LaCl}_3$ , ethanol, and para-nitrophenol. When the limiting reactant is not apparent, we can determine which reactant is limiting by comparing the molar amounts of the reactants with their coefficients in the balanced chemical equation, just as we did in Section 11.4 .

## ~~Chapter 12.2: Stoichiometry of Reactions in Solution ...~~

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